

Key word	Definition
boil (boiling)	The change of state from liquid to gas that occurs when bubbles of the substance in its gas state form throughout the liquid. Boiling occurs at the boiling point of a substance.
boiling point	The temperature at which a substance boils.
change of state	The process by which a substance changes from one state to another.
chromatogram	An image obtained from chromatography.
chromatography	A technique to separate mixtures of liquids (often coloured) that are soluble in the same solvent.
condense (condensation)	The change of state from gas to liquid. It can happen at any temperature below the boiling point.
density	The mass of a material in a certain volume.
diffusion	The process by which particles in liquids or gases spread out through random movement from a region where there are many particles to one where there are fewer.
dissolve	The complete mixing of a solute with a solvent to make a solution.
distillation	A technique that uses evaporation and condensation to obtain a solvent from a solution.
evaporate (evaporation)	The change of state from liquid to gas that occurs when particles leave the surface of the liquid only. It can happen at any temperature. Evaporation can be used to separate a solid dissolved in a liquid.
filtrate	The liquid or solution that collects in the container after the mixture has passed through the filter paper.
filtration	A way of separating pieces of solid that are mixed with a liquid or solution by pouring through filter paper.
freeze (freezing)	The change of state from liquid to solid at the melting point of a substance.
gas	In the gas state, a substance can flow and can also be compressed.

gas pressure	The force exerted per unit area on the walls of a container. It is caused by collisions of particles with the walls.
liquid	In the liquid state, a substance can flow but cannot be compressed.
material	The different types of stuff that things are made from.
melt (melting)	The change of state from solid to liquid at the melting point of a substance.
melting point	The temperature at which a substance melts.
mixture	A mixture is made up of two or more pure substances that are mixed (not chemically joined) together. A mixture's properties are different from the properties of the individual substances that make it up.
particle	A very tiny object, such as an atom or molecule, that materials are made from. They are too small to be seen with a microscope.
particle model	A way to think about how substances behave in terms of small, moving particles.
property	A quality of a substance or material that describes its appearance, or how it behaves.
pure substance	A single material with no other substances mixed with it.
residue	The solid that collects in the filter paper during filtration.
saturated solution	A solution in which no more solute can dissolve.
solid	In the solid state, a substance cannot be compressed and it cannot flow.
solubility	The maximum mass of solute that dissolves in a certain volume or mass of solvent.
solubility curve	A graph showing the change in solubility of a substance with temperature.
soluble (insoluble)	A soluble substance can dissolve in a given solvent. An insoluble substance cannot dissolve in a given solvent.

solute	The solid or gas that is dissolved in a liquid.
solution	A mixture of a solute dissolved in a solvent. All parts of the mixture are the same.
solvent	A substance, normally a liquid, that dissolves another substance.
states of matter	The three forms in which a substance can exist – solid, liquid, and gas.
sublime (sublimation)	The change of state from solid directly to gas.
substance	A material that is not a mixture. It has the same properties all the way through.